Question Bank (G scheme)

Name of subject:APPLIED CHEMISTRY Unit Test :I
Subject code: 17103 Course : ALL
Semester: I

Chapter--1.Chemical bonding (Marks-12)

(2-marks question)

1)Write composition of moder n atom.

2)Why atom is electrically neutal.

3)Arrange the following orbitals in the increasing order of energy 2p,4p,3s,3d,4s ,1s 3p. 2s.

4)Define -Isotope&Isobar.

5)Define-Atomic number&Mass number

6)Define-Valency&electrovalency

(3-marks question)

6)Write any three differences between electrovalent and covalent substances.

7)Describe the formation of water molecule(H2O)with the help of schematic diagram.

8)Describe the formation of MgO molecule.

9)Diffentiate between orbit &orbital

10)Explain the formation of NITROGEN molecule.

(4-marks question)

11)Write orbital electronic configuration of elements

7 N , 11 Na, 17Cl, 24 Cr,

12)write any four assumption of Bohr,s atomic theory.

13)writesymbol,location, charge, &mass of proton, electron and neutron.

Chapter-2.Electrochemistry (Marks-14)

(2-marks question)

14)Write two differences between strong and weak electrolytes.

15)What is reduction potential.

16)Why is Blister copper electro -refined.

17)Defineconductor&nonconductor

(3-marks questions)

18)Samequantity of current is passed through CuSO4 and Ag NO3 solution . If3.027g of Cu Is deposited, calculate weight of silver deposited(Eq -wt.of Cu=31.77,Eq –wtof Ag=108)

19)StateFaraday,sfirst law of electrolysis. Drive its equation.

20)Define PH. Calculate the ph of 0.0034N HCL assuming complete dissociation.

(**4-marks question)**

21)State and explain Electroplating of silver.

22)Describe mechanism of electrolysis of CuSO4 using platinum electrodes with the help oflabbled diagram and chemical-reaction occurring at cathode &anode.

23)Define degree of ionization .state & explain factors affecting on degree of ionization.

24)State and explain electrorefining of copper.